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HIGH QUALITY TUITION CENTRE MAGOGONI ZANZIBAR.

Example showing how to carry out a systematic qualitative analysis from a give sample of salt.

EXPERIMENT	OBSERVATION	INFERENCE
The analysis will base on salt S.		
Appearance of the salt S.	Blue crystals with no smell.	Suggest Cu^{2+} , SO_4^{2-} or SO_3^{2-} .
Solubility of the salt S in water.	Salt S soluble in cold water and form clear solution.	Suggest Cl^- , NO_3^- , SO_4^{2-} , HCO_3^- or certain SO_3^{2-} .
Action of heat on salt S.	Water vapor was given off. The salt changed from blue to white after heating.	Suggest hydrated salt of CuSO_4 . Present.
Salt S with dilute Hydrochloric acid.	No gas evolved on cold and warmed dilute HCl but the salt S dissolved and gave clear solution. The solution form white precipitates with Barium Chloride solution.	Suggest SO_4^{2-} Present.
Salt S with concentrated sulfuric acid.	No gas evolved on cold and warmed conc. Sulfuric acid but the salt changed its color to white.	Suggest SO_4^{2-} present.

CONFIRMATORY TEST FOR ANIONS

EXPERIMENT	OBSERVATION	INFERENCE
Salt S solution with Barium chloride solution.	White precipitates formed insoluble in excess dilute Hydrochloric acid.	Confirm SO_4^{2-} present.
Salt S with Sodium Hydroxide solution.	Pale blue precipitates insoluble in excess alkalis.	Confirm Cu^{2+} present.
Salt S solution with Ammonium Hydroxide solution.	Pale blue precipitates soluble in excess Ammonium Hydroxide to form deep blue solution.	Confirm Cu^{2+} present.

CONFIRMATORY TEST FOR CATIONS

EXPERIMENT	OBSERVATION	INFERENCE
Salt S solution with Potassium ferrioxalate solution.	Brown precipitates formed.	Confirm Cu^{2+} present.

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